

Drawing polyatomic ions worksheet

Continue

1. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR	11. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR	1. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR	1. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR
11. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR	11. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR	11. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR	11. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup> OR
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**Polyatomic Ions Chart**

Formula	Name	Formula	Name
NH <sub>4</sub> <sup>+</sup>	Ammonium	CrO <sub>4</sub> <sup>2-</sup>	Chromate
NH <sub>3</sub>	Ammonia	Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	Dichromate
C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> <sup>-</sup>	Acetate	MnO <sub>4</sub> <sup>-</sup>	Permanganate
CH <sub>3</sub> COO <sup>-</sup>	Acetate	MnO <sub>4</sub> <sup>2-</sup>	Manganate
CN <sup>-</sup>	Cyanide	NO <sub>2</sub> <sup>-</sup>	Nitrite
CO <sub>3</sub> <sup>2-</sup>	Carbonate	NO <sub>3</sub> <sup>-</sup>	Nitrate
HCO <sub>3</sub> <sup>-</sup>	Bicarbonate	OH <sup>-</sup>	Hydroxide
C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	Oxalate	PO <sub>4</sub> <sup>3-</sup>	Phosphate
ClO <sup>-</sup>	Hypochlorite	SCN <sup>-</sup>	Thiocyanate
ClO <sub>2</sub> <sup>-</sup>	Chlorite	Fe(CN) <sub>6</sub> <sup>3-</sup>	Ferricyanide
ClO <sub>3</sub> <sup>-</sup>	Chlorate	SO <sub>3</sub> <sup>2-</sup>	Sulfite
ClO <sub>4</sub> <sup>-</sup>	Perchlorate	SO <sub>4</sub> <sup>2-</sup>	Sulfate
S <sub>2</sub> O <sub>3</sub> <sup>2-</sup>	Thiosulfate	HSO <sub>4</sub> <sup>-</sup>	Hydrogen sulfate
BrO <sup>-</sup>	Hypobromite	IO <sub>3</sub> <sup>-</sup>	Iodate
AsO <sub>2</sub> <sup>3-</sup>	Arsenite	SeO <sub>4</sub> <sup>2-</sup>	Selenate
BrO <sub>3</sub> <sup>-</sup>	Bromate	HSO <sub>3</sub> <sup>-</sup>	Hydrogen sulfite

**Commonly Used Multivalent Metals**

Symbol	Name	-ous	-ic
Fe ( <i>Ferrum</i> )	Iron	+2	+3
Pb ( <i>Plumbum</i> )	Lead	+2	+4
Sn ( <i>Stannum</i> )	Tin	+2	+4
Hg ( <i>Hydrargyrum</i> )	Mercury	+1	+2
Cu ( <i>Cuprum</i> )	Copper	+1	+2

**Prefixes**

Mono = one	Hexa = six
Di = two	Hepta = seven
Tri = three	Octa = eight
Tetra = four	Nona = nine
Penta = five	Deca = ten

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NH <sub>4</sub> <sup>+</sup>	Ammonium	CrO <sub>4</sub> <sup>2-</sup>	Chromate
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CH <sub>3</sub> COO <sup>-</sup>	Acetate	MnO <sub>4</sub> <sup>2-</sup>	Manganate
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CO <sub>3</sub> <sup>2-</sup>	Carbonate	NO <sub>3</sub> <sup>-</sup>	Nitrate
HCO <sub>3</sub> <sup>-</sup>	Bicarbonate	OH <sup>-</sup>	Hydroxide
C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	Oxalate	PO <sub>4</sub> <sup>3-</sup>	Phosphate
ClO <sup>-</sup>	Hypochlorite	SCN <sup>-</sup>	Thiocyanate
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**Prefixes**

Mono = one	Hexa = six
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**FORMULAS WITH POLYATOMIC IONS** Name \_\_\_\_\_

Matching the horizontal and vertical axes, write the formulas of the compounds with the following combination of ions. The first one is done for you.

	OH <sup>-</sup>	NO <sub>3</sub> <sup>-</sup>	CO <sub>3</sub> <sup>2-</sup>	SO <sub>4</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>
H <sup>+</sup>	HOH (H <sub>2</sub> O)	HNO <sub>3</sub>	H <sub>2</sub> CO <sub>3</sub>	H <sub>2</sub> SO <sub>4</sub>	H <sub>3</sub> PO <sub>4</sub>
Na <sup>+</sup>					
Mg <sup>2+</sup>					
NH <sub>4</sub> <sup>+</sup>					
Ca <sup>2+</sup>					
K <sup>+</sup>					
Al <sup>3+</sup>					
Pb <sup>4+</sup>					

Macromolecules Worksheet Name: \_\_\_\_\_ Year: \_\_\_\_\_

Part A. Classify each as a carbohydrate, protein, lipid, or nucleic acid.

1. _____ starch	8. _____ DNA
2. _____ nucleotide	9. _____ sugar
3. _____ RNA	10. _____ oil
4. _____ unsaturated fat	11. _____ saturated fat
5. _____ amino acid	12. _____ meat
6. _____ enzyme	13. _____ monosaccharides
7. _____ wax	14. _____ phospholipids

Part B. Identify the specific molecule (use the above terms) from each description. Some terms may be used more than once.

- \_\_\_\_\_ monomer of proteins
- \_\_\_\_\_ genetic material
- \_\_\_\_\_ monomer of nucleic acids
- \_\_\_\_\_ one sugar
- \_\_\_\_\_ has a bent polypeptide chain
- \_\_\_\_\_ monomer of lipids
- \_\_\_\_\_ an example of a protein
- \_\_\_\_\_ monomer of carbohydrates

Part C. Which specific molecule is each food made of? (lipid, protein, carbohydrate)

1. _____ almond	7. _____ bacon
2. _____ beef jerky	8. _____ egg white
3. _____ noodles	9. _____ table sugar
4. _____ orange juice	10. _____ popcorn
5. _____ cheese	11. _____ celery
6. _____ wheat	12. _____ soy beans

Water polyatomic ions.

In order to continue enjoying our site, we ask that you confirm your identity as a human. Thank you very much for your cooperation. 1 Lewis Diagrams for Polyatomic Ions 2 Polyatomic Ions Polyatomic ions are covalently bonded atoms that together have an overall charge. 3 Lewis diagrams for polyatomic ions: Determine the # of valence e- for each atom in the molecule and draw the dot diagram as if it were covalent BUT SUBTRACT one electron (usually from the central atom) for each +ve charge. ADD one electron (usually from the central atom) for each -ve charge. 4 Draw the Lewis structure for the following: CHO21- Determine the central atom (the one with the most bonding electrons: Carbon 4 e- Hydrogen 1 e- Oxygen 2 e- per atom So the central atom will be? C 5 Draw the Lewis structure for the following: CHO21- 6 Draw the Lewis structure for the following: CHO21- H C Don't forget to add the -1 charge 7 Draw the Lewis structure for the following: CHO21- H C O O 8 Draw the Lewis structure for the following: CHO21- H C O O 9 Octet rule satisfied for every atom Draw the Lewis structure for the following: CHO21- 1- H C O O Octet rule satisfied for every atom Don't forget the brackets and the charge!!! 10 NH4+ Determine the central atom: 11 NH4+ N Remember to take one away for the +1 charge 12 NH4+ H H N H 13 NH4+ + H H N H 14 CO32- 15 CO32- C 16 CO32- C Add in the two extra electrons!!!! 17 CO32- C O 18 CO32- O C O 19 Don't forget the brackets and the charge!!! CO32- 2- O C O O Don't forget the brackets and the charge!!! 20 Where do the electrons for the charge come from in a polyatomic dot diagram? They are transferred from a metal. Which means polyatomic dot diagrams involve both ionic and covalent dot diagrams. 21 Draw the dot diagram for NaCNO First draw the dot diagram for the polyatomic. Na+ CNO1- 22 CNO1- 23 CNO1- 1- O C N 24 CNO1- 1- O C N Now finish the ionic dot diagram by drawing the dot diagram for the metal. 25 NaCNO 1- [ Na ] +1 O C N 26 Recall Ionic Dot Diagrams Draw the dot diagram for NaBr. 27 NaBr Na Br 28 NaBr [ Na ] +1 [ Br ] -1 This chemistry homework page is perfect for students to practice drawing electron dot diagrams or Lewis structures for polyatomic ions. One page focuses on simpler polyatomic ions and one focuses on polyatomic ions that have expanded octets. This is great for honors level learners. This page is part of my giant Chemistry Homework for a Year Bundle. Click here to read about that whole year bundle if you want to buy a whole year's worth of homework pages at a discount as compared to buying them separately. This page is also part of a Chemical Bonding Chemistry Homework Unit. Click here to read about this unit set. This is a non-editable PDF file. The PDF file contains the student page, the answer key, and my terms of use page. Check out my Chemistry Doodle Notes for some engaging resources that thoroughly explain Chemistry concepts in a versatile way! Contact Us! If you have any questions or concerns, please reach out to us on the question and answer section of my store and we will get back to you quickly! Terms of Use: Purchasing my teaching resources allows you to: \* make copies for your own classes only. \* place this file on your own password-protected class page or server (Blackboard, Google Drive, etc) AS LONG AS no other teacher has access to that class webpage. This resource is for you, the purchaser, alone. You are not allowed to distribute this digital resource to other teachers or post this resource on any webpage or server that is available for public view. If you and a team of teachers would like to use this resource together, please purchase additional licenses on the resource purchase page. Failure to comply with these terms of use is a copyright infringement and a violation of the Digital Millennium Copyright Act (DMCA). Clipart and elements found in this PDF are copyrighted and cannot be extracted and used outside of this file without permission or license. Files are partially or fully non-editable to protect the images that are copyrighted and purchased through licenses. Thanks for understanding! © Bethany Lau All Rights Reserved.

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